Editorial note

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Editorial note on Benzene as a good Solvent

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The ideal solvent would be non-toxic, non-flammable but combustible, radiation resistant, inert to nitric acid, easy to distil, and inexpensive. It will also need to be able to dissolve a wide variety of substances.

Since the pi cloud of the benzene ring will bind with the pi clouds of different solutes, it is more likely to dissolve several organic compounds than a saturated hydrocarbon like hexane. For this cause, benzene can dissolve polystyrene more easily than hexane. Because of its ability to dissolve stuff, benzene is considered a strong solvent. It's a decent solvent for dissolving objects, but it's notoriously difficult to deal with in the lab. The use of benzene as an NMR solvent is unique.

Although benzene is a hated solvent due to its health hazards and flammability, there are numerous aromatic solvents with better health and fire profiles. Exxon's Solves so 150ND, for example, is aromatic kerosene with different solvent properties than Statoil's "solvent 70," aliphatic saturated kerosene.

If one wants to dilute the ionic liquid aliquat 336 to 30% in a diluent and then use this solution to remove metals like copper, cobalt, and iron from aqueous chloride media, this can be beneficial. It's important to remember that not all alkanes or aromatics are the same; hexane is a common chemistry solvent, but it's also extremely toxic. It is converted to hexane2.5-dione in the body, which is extremely dangerous. The other longer alkanes, on the other hand, do not have the same toxicity. You should also think about the exposure path.

A puddle of diesel fuel on the garage floor, for example, is much less dangerous than exposing a person's hand to a high-pressure diesel spray that injects the fuel into the person's body.