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The Solid State

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Abstract

As you recognize there are three states of matter – solids, liquids, and gases. Most objects around us are in their solid state. Everything from the chair you're seated on to the sand on the beach is solid. However, their physical appearances can vary their physical properties are similar. allow us to learn a touch more.

Keywords: Solids; Liquids; Gases

- **Thermal Energy:** Thermal energy may be a sort of K.E. it's the energy that particles possess thanks to their motion. It's the interior energy of an object that's liable for its temperature. Transfer of thermal energy happens through the transfer of warmth. With an increase within the thermal energy, the particles of matter tend to manoeuvre faster and vice-versa.

These two forces keep the atoms of solids closely packed during a fixed structure we call a lattice. At low temperatures the thermal energy of particles of solid matter is low. Therefore the movement of particles is minimal, the intermolecular force is high and spaces between atoms are extremely less. This provides solid their basic properties.

Introduction

As we've studied before we will change the physical properties of an object or element by varying its temperature and pressure, this phenomenon is understood as change of State. For instance once we heat water to show it into water vapour or freeze it to show it into ice.

The states of matter and its stability all depends on two opposing forces found within the atom of a component. These are the two forces liable for the formation and stability of a solid. Allow us to discuss these two forces in some detail.

- **Intermolecular Forces:** Intermolecular forces exist between molecules of a component. They really determine not only the state of the matter but also factors like boiling point, freezing point, enthalpy of the weather. They keep the atoms or molecules of matter approximate during a bond.

Conclusion

The Combinatorial Chemistry may be a methodology during which a really sizable amount of chemical entities are synthesized by condensing a little number of chemical compounds together altogether combinations defined by a little set of chemical reactions.

References

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